

THE KENYA NATIONAL EXAMINATIONS COUNCIL  
Kenya Certificate of Secondary Education

233/1 -

**CHEMISTRY**  
**(THEORY)**

- Paper 1

Nov. 2018 - 2 hours

Name ..... Index Number .....

Candidate's Signature ..... Date .....

**Instructions to candidates**

- Write your name and index number in the spaces provided above.
- Sign and write the date of examination in the spaces provided above.
- Answer **ALL** the questions in the spaces provided in the question paper.
- KNEC mathematical tables and silent non-programmable electronic calculators may be used.
- All working **MUST** be clearly shown where necessary.
- This paper consists of 16 printed pages.**
- Candidates should check the question paper to ascertain that all the pages are printed as indicated and that no questions are missing.
- Candidates should answer the questions in English.



**For Examiner's Use Only**

1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16
17	18	19	20	21	22	23	24	25	26	27	28	Grand Total			

1. (a) Define a soluble base. (1 mark)

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- (b) Aqueous solutions of 2M ethanoic acid and 2M nitric(V) acid were tested for electrical conductivity. Which solution is a better conductor of electricity? Explain. (2 marks)

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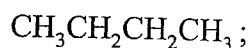
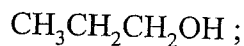
2. (a) Explain why it is not advisable to prepare a sample of carbon(IV) oxide using barium carbonate and dilute sulphuric(VI) acid. (2 marks)

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- (b) State a method that can be used to collect dry carbon(IV) oxide gas. Give a reason. (1 mark)

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3. The following are formulae of organic compounds. Use the formulae to answer the questions that follow:



- (a) Select:

- (i) **two** compounds which when reacted together produce a sweet smelling compound. (1 mark)

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- (ii) an unsaturated hydrocarbon. (1 mark)

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- (b) Name the compound selected in (a) (ii). (1 mark)

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4. One of the allotropes of sulphur is rhombic sulphur.

- (a) Name the other allotrope of sulphur. (1 mark)

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- (b) Draw a diagram to show the shape of the allotrope named in (a) above. (1 mark)

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(c) Write an equation for the reaction between concentrated sulphuric(VI) acid and sulphur. (1 mark)

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5. Describe an experiment to show that group one elements react with cold water to form alkaline solutions. (3 marks)

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6. (a) State Graham's law of diffusion. (1 mark)

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(b) Explain why a balloon filled with helium gas deflates faster than a balloon of the same size filled with argon gas. (2 marks)

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7. 30.0 cm<sup>3</sup> of aqueous sodium hydroxide containing 8.0 g per litre of sodium hydroxide were completely neutralised by 0.294 g of a dibasic acid. Determine the relative formula mass of the dibasic acid. (Na = 23.0 ; O = 16.0 ; H = 1.0) (3 marks)

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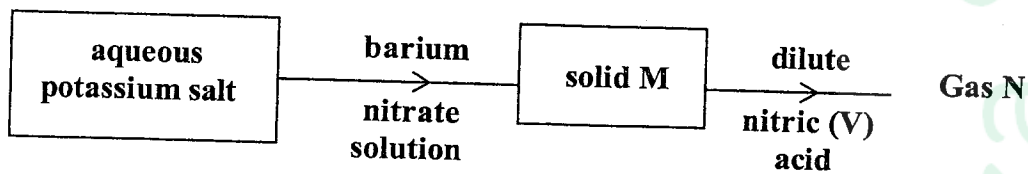
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8. Study the flow chart in **Figure 1** and answer the questions that follow.



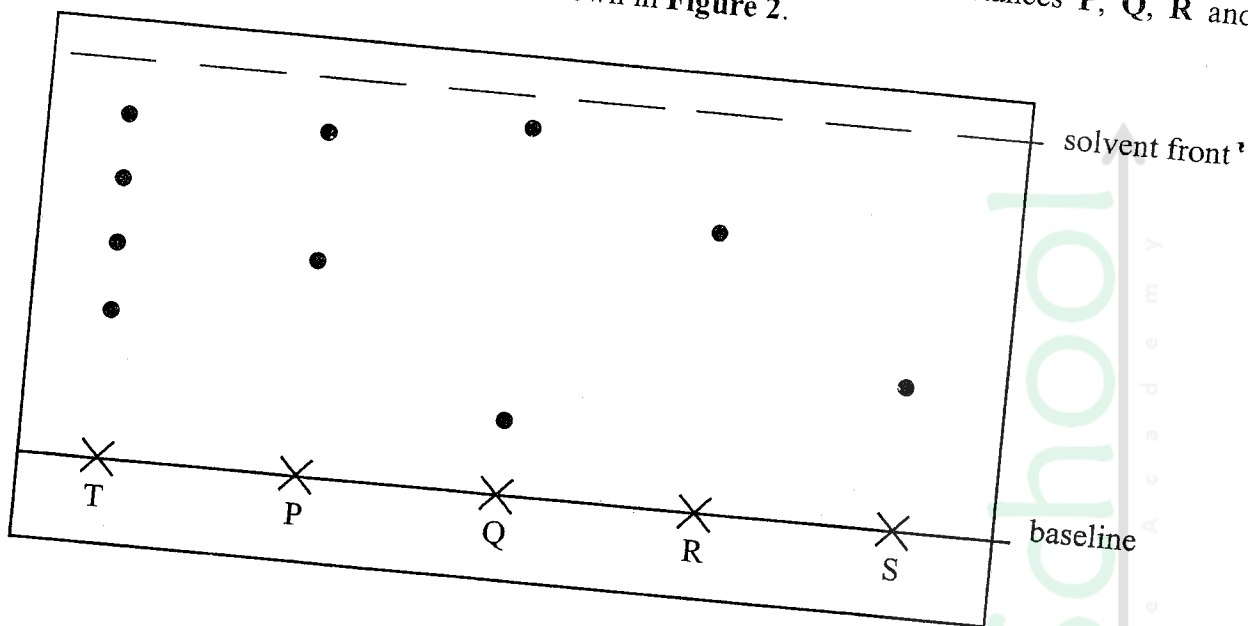
**Figure 1**

Gas N forms a white suspension with aqueous calcium hydroxide.

- (a) Name the anion present in the potassium salt. (1 mark)
- .....
- (b) Write an ionic equation for the formation of solid M. (1 mark)
- .....
- .....
- (c) Give **one** use of gas N. (1 mark)
- .....



9. An experiment was carried out to determine the presence of substances P, Q, R and S in mixture T. The results obtained are shown in **Figure 2**.



**Figure 2**

- (a) Name the method of separation illustrated in **Figure 2**. (1 mark)
- .....
- (b) Select:
- (i) one substance which contains a component **not** present in T. (1 mark)
- .....
- (ii) a substance which is least soluble in the solvent used. (1 mark)
- .....

10. Using iron filings, describe an experiment that can be conducted to show that oxygen is present in air. (3 marks)

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11. (a) Element U has atomic number 12 while element V has atomic number 16. How do the melting points of their oxides compare? Explain. (3 marks)

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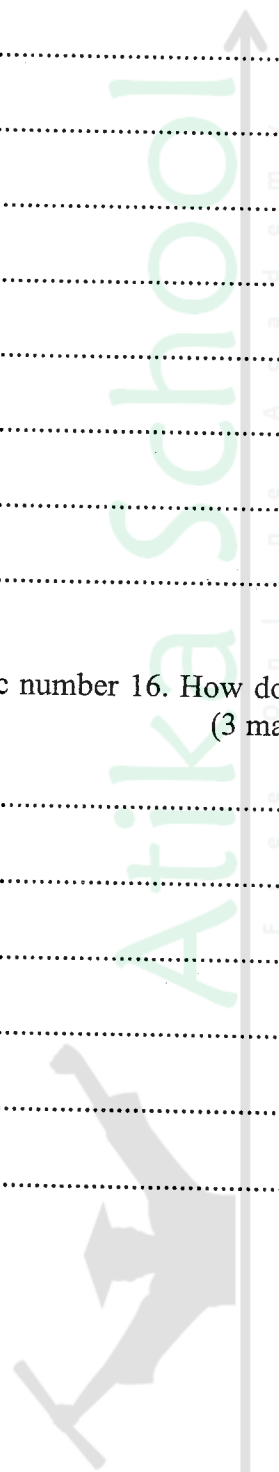
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12. When ethene gas is compressed at a high temperature, a solid is formed.

(a) Give the name of the solid.

(1 mark)

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(b) Explain why it is **not** advisable to allow the solid to accumulate in the environment.

(2 marks)

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13. In the Haber process, nitrogen reacts with hydrogen according to the following equation.



(a) What would be the effect of adding a catalyst on the position of the equilibrium?

(1 mark)

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(b) Explain why it is **not** advisable to use temperatures higher than 773 K in the Haber process.

(2 marks)

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14. Figure 3 shows a set-up used by a student to prepare dry chlorine gas in the laboratory.

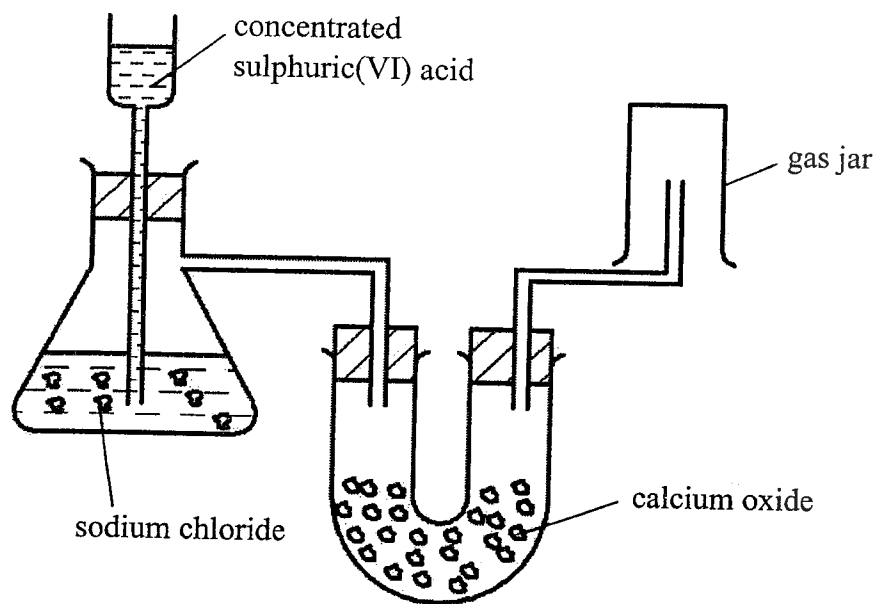


Figure 3

Identify **three** mistakes in the set-up, and give a reason for each.

(3 marks)

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15. You are provided with solid potassium hydrogen carbonate. Describe how a solid sample of potassium nitrate can be prepared. (3 marks)

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16. Metal X and Y have standard electrode potentials of  $-0.13\text{ V}$  and  $-0.76\text{ V}$  respectively. The metals were connected to form a cell as shown in Figure 4.

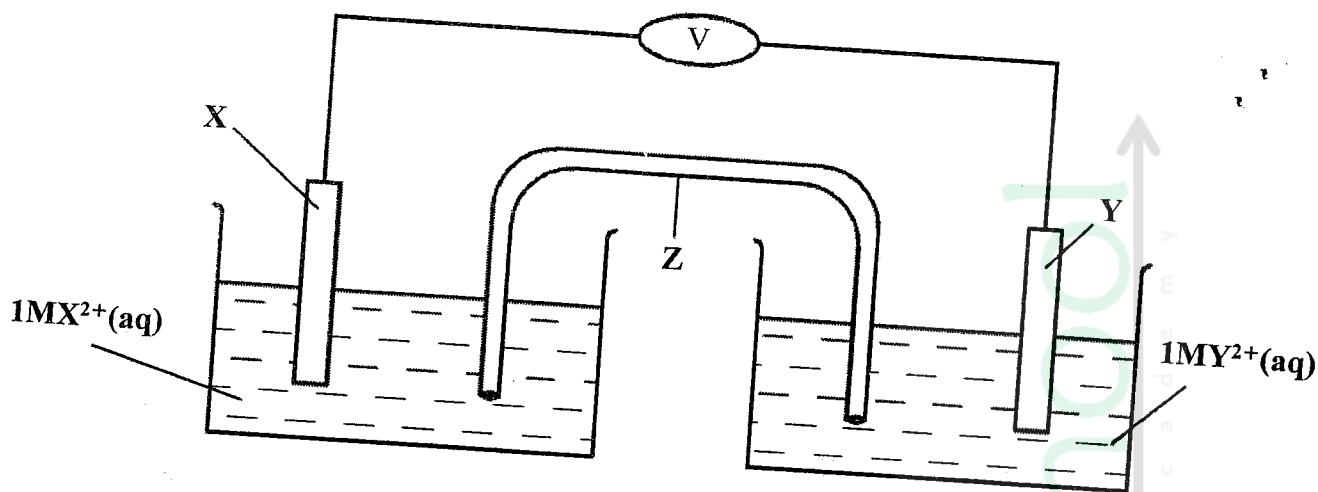


Figure 4

- (a) Name the part labelled Z. (1 mark)
- .....
- (b) State **one** function of the part labelled Z. (1 mark)
- .....
- .....
- .....
- .....
- (c) Calculate the e.m.f. of the cell. (1 mark)
- .....
- .....
- .....

17. **Figure 5** represents a grid that is part of the periodic table. Study it and answer the questions that follow. The letters are not the actual symbols of the elements.

	A				B			
				C				
	E							

**Figure 5**

- (a) Write the electron arrangement of element C. (1 mark)
- .....
- (b) On the grid provided, show with a tick (✓) the position of element D whose atomic number is 18. (1 mark)
- (c) Element E is more reactive than A. Explain. (1 mark)

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18. (a) Define molar heat of displacement. (1 mark)

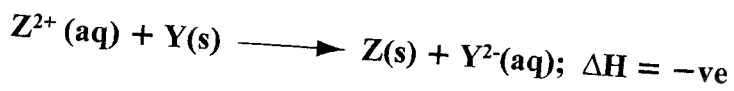
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- (b) The following ionic equation represents the reaction between metal Y and an aqueous solution of  $Z^{2+}$ .



Draw an energy level diagram to represent the reaction.

(2 marks)

19. (a) Give the symbols of the two charged particles emitted by a radioactive isotope.

(1 mark)

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- (b) An isotope  $^{210}_{82}\text{Pb}$  disintegrates by emitting two beta particles. Determine the mass number and atomic number of the resulting nuclide.

mass number:

(1 mark)

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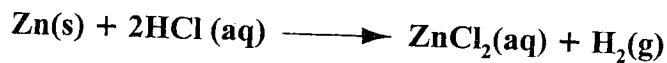
atomic number:

(1 mark)

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20. (a) Zinc reacts with hydrochloric acid according to the following equation.



Identify the reducing agent. Give a reason for the answer.

(2 marks)

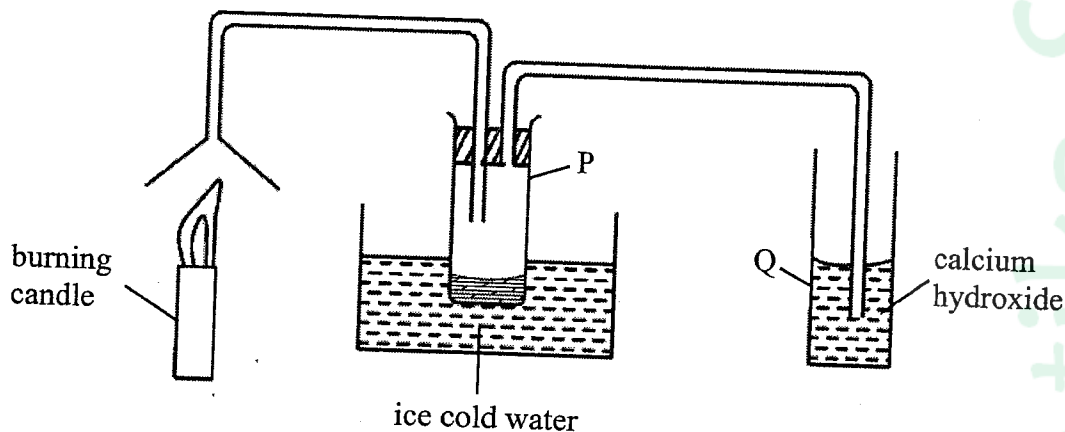
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- (b) Iron sheets are dipped in molten zinc to prevent rusting. Name this process. (1 mark)

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21. Study the set-up in **Figure 6** and answer the questions that follow.



**Figure 6**

- (a) Name the substance that was collected in tube P. (1 mark)

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- (b) Write an equation for the reaction which occurs in tube Q in the first few minutes of the experiment. (1 mark)

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- (c) Give a suitable conclusion for the experiment in the set-up. (1 mark)

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22. You are provided with the following: thermometer, boiling tube, beaker, Bunsen burner, pure substance **X** whose boiling point is about  $80^{\circ}\text{C}$ , water and any other apparatus that may be required. Draw a labelled diagram of the set-up that can be used to determine the melting point of **X**. (3 marks)

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23. Explain why it is important to put off a non-luminous flame immediately after use. (2 marks)

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24. (a) Name **two** ores of iron. (1 mark)

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- (b) Describe how the amount of iron in a sample of iron(III) oxide can be determined. (2 marks)

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25. Explain why a solution of sodium chloride conducts electricity while that of sugar does not. (2 marks)

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26. Explain why commercial indicators are preferred to flower extracts as acid-base indicators. (2 marks)

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27.  $(\text{NH}_4)_2\text{HPO}_4$  is a fertiliser used by farmers to boost their crop production.

(a) Calculate the mass of phosphorus in a 20 kg packet of  $(\text{NH}_4)_2\text{HPO}_4$ .

(N = 14.0; H = 1.0; P = 31.0; O = 16.0)

(2 marks)

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(b) State **one** advantage of this fertilizer,  $(\text{NH}_4)_2\text{HPO}_4$ , over urea  $(\text{CO}(\text{NH}_2)_2)$ . (1 mark)

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28. Distinguish between empirical and molecular formula of a compound.

(2 marks)

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